

Bohr Models

Name: *Key*
 Date:
 Block:

Bohr models show all electrons in each electron shell.

Electrons are distributed in the following way:

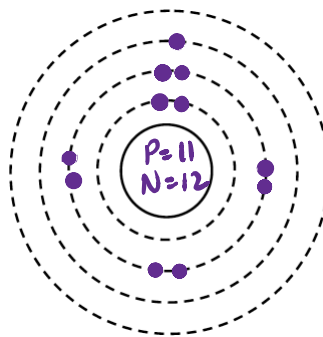
- The first electron shell holds a maximum of 2 electrons.
- Each of the next shells holds a maximum of 8 electrons.
- Shells cannot be created until the lower shell is completely filled.

In the Bohr model diagram, make sure to:

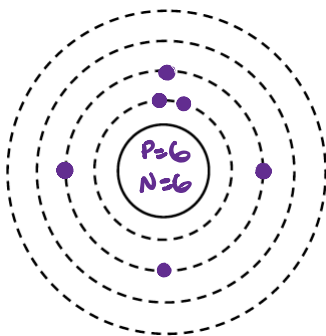
1. Write the number of protons and neutrons in the center of the diagram.
2. Place electrons one at a time before they are paired up.

Example: Sodium

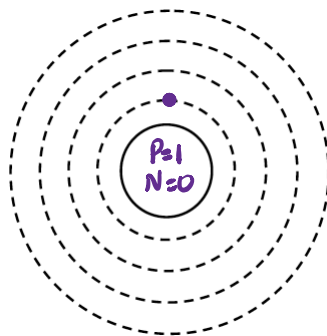
Atomic # = 11
 Atomic mass = 23
 Protons = 11
 Electrons = 11
 Neutrons = $23 - 11 = 12$



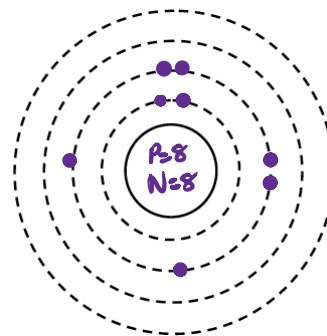
1. Carbon



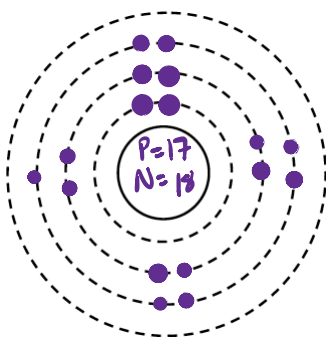
2. Hydrogen



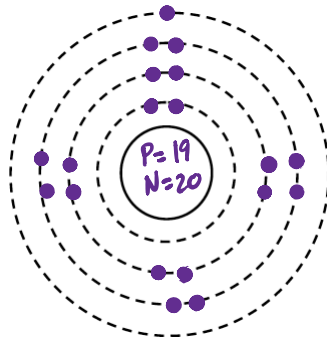
3. Oxygen



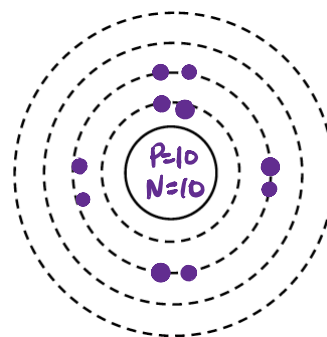
4. Chlorine



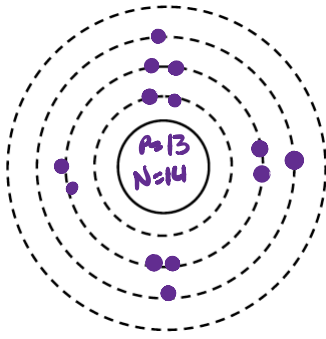
5. Potassium



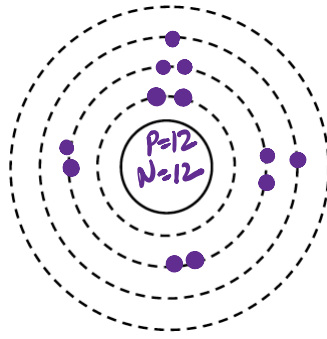
6. Neon



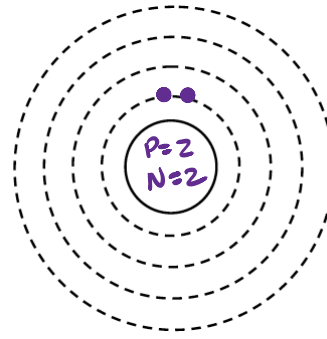
7. Aluminum



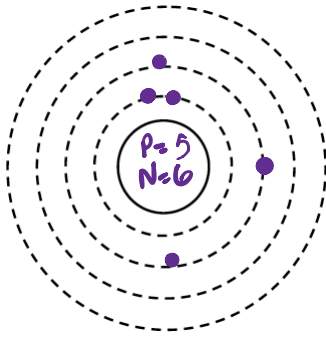
8. Magnesium



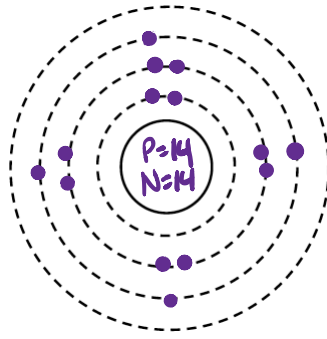
9. Helium



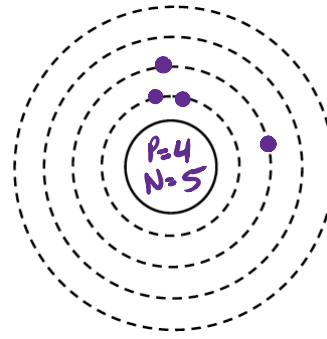
10. Boron



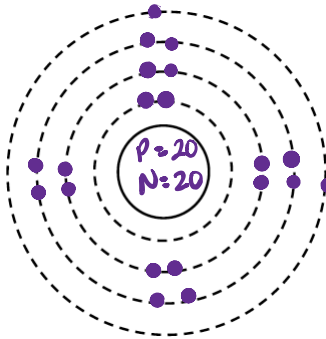
11. Silicon



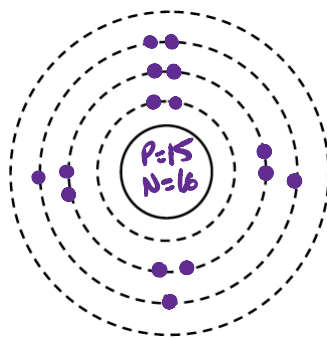
12. Beryllium



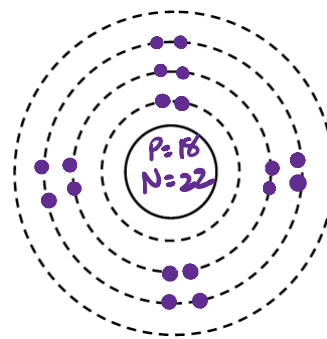
13. Calcium



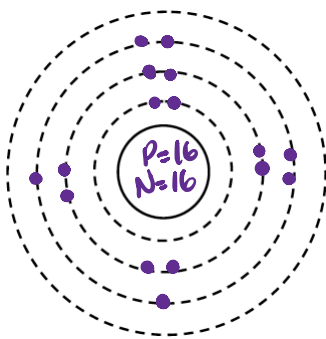
14. Phosphorus



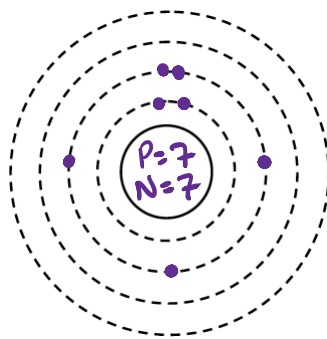
15. Argon



16. Sulfur



17. Nitrogen



18. Fluorine

