

Chemistry 11
Science 10 Review

Name: *Key*
Date:
Block:

Write the formula for the following compounds:

1. Barium nitride
 Ba_3N_2
2. Lithium oxide
 Li_2O
3. Aluminum sulphide
 Al_2S_3
4. Copper(II)oxide
 CuO
5. Zinc(II) fluoride
 ZnF_2
6. Chromium(III) nitride
 CrN
7. Silver oxalate
 $\text{Ag}_2\text{C}_2\text{O}_4$
8. Chromium(II) acetate
 $\text{Cr}(\text{CH}_3\text{COO})_2$
9. Gold (III) chromate
 $\text{Au}_2(\text{CrO}_4)_3$

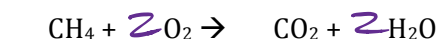
Write the name for the following compounds:

1. GaAs
Gallium arsenide
2. KCl
Potassium chloride
3. Ni_3P_2
Nickel (II) phosphide
4. TiO_2
Titanium (IV) oxide
5. Au_2Se
Gold (I) selenide
6. $\text{Ag}_2\text{C}_2\text{O}_4$
Silver (I) dichromate
7. Na_2CrO_4
Sodium chromate
8. $\text{Cr}(\text{NO}_3)_2$
Chromium (II) nitrate

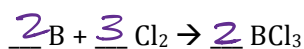
Classify each of the following as an acid, base, or salt.

1. Potassium hydroxide base
2. Has a pH of 3 acid
3. NaCl salt
4. Magnesium carbonate salt
5. HF acid
6. LiOH base
7. Tastes bitter base
8. Calcium nitrate salt
9. H₃PO₄ acid
10. Sr(OH)₂ base

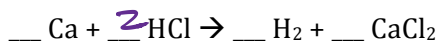
Balance the following reaction and classify the type of chemical reaction:



Combustion



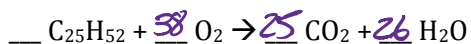
Synthesis



Single replacement



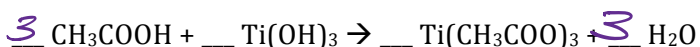
Synthesis



Combustion



decomposition



neutralization

Write the complete balanced chemical reactions for the following:

- a) Potassium hydroxide and hydrogen are produced when potassium metal reacts with water.



- b) The reaction between magnesium metal and copper(II) sulphate.



- c) Decomposition of mercury(II) oxide to its elements.

